



Across

1. charge of ion made by oxygen is negative _____
3. The unique properties of water are due to _____ bonding.
4. The type of molecules that dissolve in oil _____
7. Most solids dissolve more in _____ water
9. Element with electron config $1s^2 2s^2 2p^4$
10. Electron sharing is _____ bonding.
11. Standard Temperature is _____ °C
12. Gas _____ is related to number of collisions
15. NaOH would make a _____ solution
20. Metals want to _____ electrons to become octets
21. Most gases dissolve more in _____ water
23. The weakest force, holds non-polar molecules together
25. pH of a solution with $[\text{OH}^-] = 10^{-8} \text{ M}$
26. Number of outer electrons (dots) in flourine
27. Shape of a water molecule
28. The _____ acid of H_2O is H_3O^+
30. The pH of a weak base can be about _____.
31. As the temperature increases, the particles move _____

DOWN CLUES

1. Shape of molecule with 4 bonding electron pairs
2. As the temperature increases, the pressure of a gas _____
5. Standard Pressure is _____ atm.
6. A(n) _____ compound is a brittle non-conductor.
8. A molecule like H_2O with + and - ends.
13. A(n) _____ solution contains the maximum amount of dissolved solute
14. F^- ion has the same electron config. as _____
16. A(n) _____ makes a solution with lots of H^+ ions.
17. Element in 3rd series that wants to *gain one* electron
18. For gas law problems, the temperature must be in _____
19. The _____ the mass of the non-polar molecule, the weaker the force.
22. Element with 14 electrons
24. Number of dots (or electrons) in a happy atom.
29. Particles move rapidly and at random in a _____.